

## SIGNIFICANT FIGURES – TOPIC TEST 1

### QUESTION 1

The measurement, 206 cm, has how many significant (measured) digits?

- A 1
- B 2
- C 3
- D 4
- E 5

### QUESTION 2

The measurement, 206.0°C, has how many significant digits?

- A 1
- B 2
- C 3
- D 4
- E 5

### QUESTION 3

The measurement, 0.00206 g, has how many significant digits?

- A 1
- B 2
- C 3
- D 4
- E 5

### QUESTION 4

The measurement, 0.0020600 mole, has how many significant digits?

- A 1
- B 2
- C 3
- D 4
- E 5

### QUESTION 5

The measurement,  $2.060 \times 10^{-3}$  coulombs, has how many significant digits?

- A 1
- B 2
- C 3
- D 4
- E 5

**QUESTION 6**

The measurement, 20600 molecules, has how many significant digits?

- A 1
- B 2
- C 3
- D 4
- E 5

**QUESTION 7**

Add the following three numbers and report your answer using significant figures:

$$2.5 \text{ cm} + 0.50 \text{ cm} + 0.055 \text{ cm} = ?$$

- A 3.055 cm
- B 3.06 cm
- C 3.1 cm
- D 3.0 cm
- E 3 cm

**QUESTION 8**

Subtract the following numbers and report your answer using significant figures:

$$416 \text{ g} - 210 \text{ g} = ?$$

- A 206.0 g
- B 206 g
- C 210. g
- D 210 g
- E 200 g

**QUESTION 9**

Multiply the following three numbers and report your answer to the correct number of significant figures:

$$0.020 \text{ cm} \times 50 \text{ cm} \times 11.1 \text{ cm} = ?$$

- A  $10 \text{ cm}^3$
- B  $11.0 \text{ cm}^3$
- C  $11. \text{ cm}^3$
- D  $11.1 \text{ cm}^3$
- E  $11.10 \text{ cm}^3$

**QUESTION 10**

Divide the following three numbers and report your answer to the correct number of significant figures:

$$0.530 \text{ g} / 0.1010 \text{ mL} = ?$$

- A 2 g/mL
- B 5.2 g/mL
- C 5.3 g/mL
- D 5.25 g/mL
- E 5.248 g/mL

## **ANSWERS**

<b>QUESTION 1</b>	Answer is C
<b>QUESTION 2</b>	Answer is D
<b>QUESTION 3</b>	Answer is C
<b>QUESTION 4</b>	Answer is E
<b>QUESTION 5</b>	Answer is D
<b>QUESTION 6</b>	Answer is E
<b>QUESTION 7</b>	Answer is C
<b>QUESTION 8</b>	Answer is B
<b>QUESTION 9</b>	Answer is C
<b>QUESTION 10</b>	Answer is D